

Oxidative Cyclizations in a Nonpolar Solvent Using Molecular Oxygen and Studies on the Stereochemistry of Oxypalladation

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Abstract: Oxidative cyclizations of a variety of heteroatom nucleophiles onto unactivated olefins are catalyzed by palladium(II) and pyridine in the presence of molecular oxygen as the sole stoichiometric oxidant in a nonpolar solvent (toluene). Reactivity studies of a number of N-ligated palladium complexes show that chelating ligands slow the reaction. Nearly identical conditions are applicable to five different types of nucleophiles: phenols, primary alcohols, carboxylic acids, a vinylogous acid, and amides. Electron-rich phenols are excellent substrates, and multiple olefin substitution patterns are tolerated. Primary alcohols undergo oxidative cyclization without significant oxidation to the aldehyde, a fact that illustrates the range of reactivity available from various Pd(II) salts under differing conditions. Alcohols can form both fused and spirocyclic ring systems, depending on the position of the olefin relative to the tethered alcohol; the same is true of the acid derivatives. The racemic conditions served as a platform for the development of an enantioselective reaction. Experiments with stereospecifically deuterated primary alcohol substrates rule out a "Wacker-type" mechanism involving anti oxypalladation and suggest that the reaction proceeds by syn oxypalladation for both mono- and bidentate ligands. In contrast, cyclizations of deuterium-labeled carboxylic acid substrates undergo anti oxypalladation.

Introduction

Oxidation is one of the most fundamental and important processes in nature. It has also become one of the most effective ways for chemists to induce asymmetry in organic transformations for the production of enantioenriched materials.¹ Most enantioselective oxidations involve the transfer of a heteroatom, commonly oxygen, to a substrate in a manner analogous to that of oxygenase metalloenzymes. In contrast, there is a significant lack of asymmetric two-electron oxidations that do not involve heteroatom transfer. This second mode of oxidation is similar to oxidase-type enzymatic reactions for which oxygen atoms instead act as electron and proton acceptors in substrate oxidation.

The ability of palladium (Pd) to serve as both a nucleophile (Pd(0)) and an electrophile (Pd(II)) has led to an explosion in Pd(0)-catalyzed reactions that have become indispensable to organic chemistry.² This property has also led to the develop-

ment of *racemic* oxidase-type reactions that feature electrophilic Pd(II), some examples of which are shown in Figure 1, although they are much less prevalent.³ Such reactions require the recycling of Pd(0) to Pd(II) by a stoichiometric oxidant. Ideally, such a process would take advantage of the most abundant and inexpensive oxidant, molecular oxygen (O₂). In fact, one of the early systems developed for the catalytic oxidation of secondary alcohols by Pd(II) employed O₂ directly as the only stoichiometric oxidant.⁴ The use of DMSO as solvent also has enabled *racemic* O₂-coupled Pd(II)-catalyzed oxidative cyclizations of olefin-appended nucleophiles, such as phenols, alcohols, acids, and tosylamides, as well as alcohol oxidations to ketones and aldehydes.⁵ Two other reoxidation systems include the traditional Wacker process that employs a copper cocatalyst in the presence of O₂ and the use of the organic oxidant benzoquinone as a stoichiometric oxidant. *Racemic* cyclization reactions have also been developed using these reoxidants.^{6,7}

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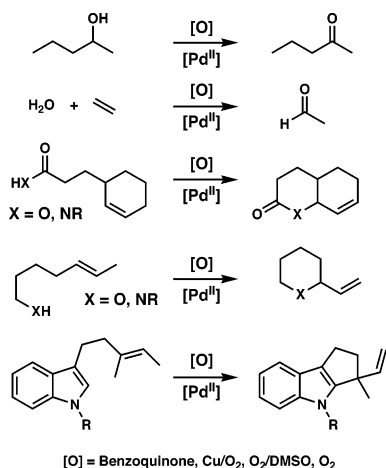


Figure 1. Prototypical Pd(II)-catalyzed dehydrogenation reactions.

Despite these advances in racemic Pd(II)-catalyzed oxidation reactions, many do not meet the ideal criteria for *enantioselective* reactions of employing only O₂ as the stoichiometric oxidant in a noncoordinating solvent. Instead, the necessary use of cocatalysts, organic oxidants, or DMSO has complicated the development of asymmetric oxidase-type reactions. For example, the use of copper/O₂ introduces a secondary catalytic cycle and another metal that could compete with Pd for the chiral ligand. The benzoquinone system requires the removal of stoichiometric amounts of organic compounds,⁸ and benzoquinone itself can act as a ligand.^{9,10} In addition, DMSO is a highly donating solvent that could interfere with the coordination of a chiral ligand to Pd. Despite the obstacles presented by the traditional oxidation systems, seminal works by Hosokawa and Murahashi,¹¹ Hayashi,¹² Sasai,¹³ and Bäckvall⁹ established the potential of enantioselective Pd(II)-catalyzed oxidative cyclizations and dialkoxylations.¹⁴

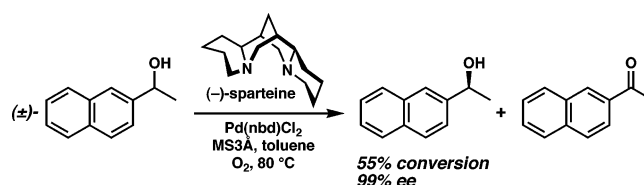


Figure 2. An oxidase-type kinetic resolution of secondary alcohols.

Table 1. Optimization of Pd Source

entry	Pd source	additive	time	yield ^b
1	Pd(nbd)Cl ₂	none	24 h	7%
2	PdCl ₂	none	24 h	27%
3	Pd(OAc) ₂	none	24 h	76%
4	Pd(TFA) ₂	none	60 min	87%
5	Pd(TFA) ₂	Na ₂ CO ₃	20 min	95%
6	Pd ₂ (dba) ₃	none	24 h	25%
7	Pd black	none	24 h	NR
8	none	none	24 h	NR
9	Pd(TFA) ₂	none, no O ₂	24 h	24%
10 ^c	Pd(TFA) ₂	Na ₂ CO ₃ , Hg ⁰	5 h	84% ^d

^a With 5 mol % of Pd source, 20 mol % of pyridine, 500 mg/mmol of MS3A, 1 atm O₂, toluene (0.1 M), 80 °C. ^b Isolated yield. ^c With 5 mol % of (pyridine)₂Pd(TFA)₂, 10 mol % of pyridine, 2 equiv of Na₂CO₃, 30 equiv of Hg⁰. ^d Conversion determined by GC.

With the requirement that O₂ be the only stoichiometric oxidant, we began to address the lack of enantioselective oxidase-type catalysis with the development of the Pd(II)-catalyzed asymmetric dehydrogenation of secondary alcohols illustrated in Figure 2.^{15,16} This oxidation effects a kinetic resolution to provide enantioenriched alcohol and was the first example of an asymmetric oxidase-type reaction in that it employs O₂ as the terminal oxidant. The basis for this chemistry was another racemic Pd(II)-catalyzed alcohol oxidation reported in 1999 by Uemura and co-workers.¹⁷ Significantly, the reaction used only O₂ in a noncoordinating solvent (toluene) and was accelerated by the presence of a ligand (pyridine). We subsequently initiated an effort to apply our enantioselective alcohol oxidation to the development of asymmetric versions of reactions such as those shown in Figure 1. In line with this

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Table 2. Oxidative Cyclizations with Substituted Pyridyl and Alkyl Amine Ligands

entry ^b	ligand	no additive	40% ligand	40% ligand + 40% Na ₂ CO ₃	entry	ligand	no additive	40% ligand	40% ligand + 40% Na ₂ CO ₃
1.		98% 1 h	99% 1 h	99% 15 min	6. ^d		92% 18 h	97% 5 h	93% 5 h
2.		85% 2.5 h	96% 9 h	99% 2 h	7. ^d		94% 8 h	96% 24 h	99% 24 h
3.		50% 24 h ^c	68% 24 h	91% 5 h	8.		82% 20 h ^c	90% 5 h ^c	92% 5 h ^c
4.		84% 8 h ^c	99% 40 min	96% 15 min	9.		15% 24 h ^c	15% 24 h ^c	10% 24 h ^c
5.		70% 30 min ^c	99% 30 min	99% 5 min	10. ^d		73% 12 h	97% 8 h	99% 8 h
					11. ^d		34% 20 h ^c	54% 20 h ^c	76% 11 h ^c

^a With 5 mol % of $L_n\text{Pd}(\text{TFA})_2$, 500 mg/mmol of MS3A, 1 atm O₂, toluene (0.1 M), 80 °C. ^b Conversion determined by GC. ^c Pd black precipitate was observed in the reaction mixture. ^d For entries 6, 7, 10, and 11, 20 mol % of excess ligand was added.

goal, we recently communicated the extension of Uemura's conditions to the cyclizations of heteroatoms and electron-rich indoles onto pendant olefins, as well as an asymmetric version of the former reaction.^{18,19} This work established that heteroatom–olefin cyclizations that use O₂ as the sole oxidant could be amenable to asymmetric catalysis. We now report the full details regarding the development of these racemic and asymmetric heterocyclizations, an expanded substrate scope for both, and deuterium-labeling studies that establish the stereochemistry of oxypalladation for two types of substrates. On the basis of our results, we offer a rationale for the difficulty in developing the asymmetric reaction in this system specifically. As recent work by Hayashi, Wolfe, Sanford, White, Stahl, Sigman, and others has also shown, our studies further demonstrate that oxidase-type catalysis by Pd(II) is both advantageously versatile in terms of reactivity and frustratingly promiscuous in terms of mechanism.

Results and Discussion

Reaction Development. The Effect of Palladium X[–] Ligand, Exogenous Base, and Nitrogen-Containing Ligand on Reaction Rate. Our initial aim was to establish Pd(II)-catalyzed conditions for racemic aerobic cyclizations to which a chiral ligand eventually could be introduced. Thus, we began our investigation of aerobic oxidative cyclizations with 2-(1,2-dimethyl-1-propenyl)phenol (**1**) using a variety of palladium(II) salts, pyridine, O₂, and MS3A in toluene at 80 °C (Table 1). These conditions are modeled after Uemura's alcohol oxidation conditions,^{17a} which, as stated above, were also employed as a starting point for our oxidative kinetic resolution chemistry.^{15a} Surprisingly, Pd(nbd)Cl₂, which is the most effective catalyst for the kinetic resolution chemistry, was ineffective for the cyclization of **1** to dihydrofuran **2** (entry 1). Treatment of **1** with a range of palladium(II) salts (entries 1–4)

led to the discovery that the electron-deficient palladium(II) trifluoroacetate (Pd(TFA)₂) was most effective for producing **2** in good yield after reasonable reaction time (entry 4). A number of exogenous bases were examined with the hope that proton consumption would accelerate the reaction (not shown, see Supporting Information). Sodium carbonate exerts the most positive effect to give 95% yield in under 30 min (entry 5). Pd(0) sources were found to be poor catalysts for the reaction: Pd₂(dba)₃ resulted in the formation of a small amount of product, along with palladium black (entry 6), and palladium black itself gave no reaction (entry 7). No cyclization occurred in the absence of any palladium (entry 8). O₂ is necessary for good reactivity, although there appears to be a background reaction that leads to cyclization since the product is formed in greater than 5% yield (entry 9). The presence of 30 equiv of elemental mercury slowed the cyclization, but did not prevent reaction altogether, which contraindicates colloidal Pd or Pd nanoparticles as the relevant catalytic species (entry 10).²⁰

Although pyridine had been a competent ligand during our optimization studies, we carried out a small ligand screen of other nitrogen-containing ligands.²¹ Each $L_n\text{Pd}(\text{TFA})_2$ complex was synthesized separately and characterized, rather than generated in situ, to limit uncertainty regarding the catalyst or catalyst precursor. Reactions were performed with no additive, 40 mol % excess ligand, and both excess ligand and Na₂CO₃. As shown in Table 2, substituted pyridyls that are less coordinating than pyridine, whether due to electronic or steric reasons, result in the precipitation of Pd black. Bidentate nitrogen-containing ligands, such as 2,2'-dipyridyl (entry 6), 4,7-dimethyl-1,10-phenanthroline (entry 7), TMEDA (entry 10), or TMPDA (entry 11), significantly slow the reaction. Weak alkylamine donors result in the precipitation of Pd black

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Table 3. Oxidative Cyclizations of Phenols^a

entry	substrate	product	time	yield	entry	substrate	product	time	yield
1.			20 min	95%	9.			10 min	86%
2.			20 min	99%	10.			2 h	93%
	R = Me 3	R = Me 4			11.			25 min	80%
3.	R = <i>t</i> -Bu 5	R = <i>t</i> -Bu 6	25 min	90%	12.			3 h	74%
4.	R = OMe 7	R = OMe 8	15 min	89%	13. ^b			75 min	85%
5.	R = COCH ₃ 9	R = COCH ₃ 10	25 h	93%	14. ^c			14 h	86%
6.	R = Br 11	R = Br 12	24 h	33%	15.		NR	24 h	
7.			20 min	85%					
	R = Me 13	R = Me 14							
8.	R = OMe 15	R = OMe 16	40 min	80%					

^a With 5 mol % of Pd(TFA)₂, 20 mol % of pyridine, 2 equiv of Na₂CO₃, 500 mg/mmol of MS3Å, 1 atm O₂, toluene (0.1 M), 80 °C. ^b The starting material was used as a 3.6:1 mixture of olefin isomers. ^c With 2 mol % of Pd(TFA)₂, 8 mol % of pyridine.

(entries 8, 9, and 11). Although some rate enhancement was observed for the nicotinate derivatives (entries 4 and 5), pyridine offered the best combination of reactivity, catalyst stability, and availability.

Cyclizations of Phenols and Primary Alcohols. Under our optimized conditions, oxidative cyclization of a variety of substituted phenols occurs readily with 5 mol % of Pd(TFA)₂, 20 mol % of pyridine, 2 equiv of Na₂CO₃, and 500 mg of 3 Å molecular sieves/mmol substrate at 0.1 M concentration in toluene under a balloon of oxygen (Table 3). Workup involves simple filtration through a pad of silica gel. Cyclizations of electron-rich phenols are facile and provide excellent yields in under 30 min (entries 1–4, 7–10). An electron-deficient phenol (**9**) serves as an excellent substrate as well, albeit with slower reaction time (entry 5). In contrast, a *p*-bromo-substituted substrate (**11**, entry 6) appears to undergo alternate reaction pathways leading to decomposition, possibly via oxidative addition to Pd(0). Cyclization onto a tetrasubstituted olefin (**21**) proceeds in good yield (entry 11), as does cyclization of a disubstituted olefin (**23**, entry 12). A dihydropyran (**26**) is available under identical conditions (entry 13). Finally, high yields and reasonable rates persist with reduced catalyst loading (2 mol %, entry 14). For terminal olefin substrate **27**, reaction does not take place, most likely because *exo* cyclization occurs that then leaves no β-hydrogens to be eliminated from the presumed Pd–alkyl intermediate.

In addition to phenols, we have investigated primary alcohol/olefin oxidative cyclizations. Remarkably, these reactions proceed to the heterocyclic ethers with, in most cases, little or no oxidation to the aldehyde under our optimized conditions (Table 4). In addition to benzyl alcohol **28**, cyclopentene (**30** and **32**) and cyclohexene (**34**) derivatives provide moderate to

Table 4. Oxidative Cyclizations of Primary Alcohols^a

entry	substrate	product	time	yield
1. ^b			3 h	87%
2.			10 h	93% ^c
3.			7.5 h	69% ^d
4.			20 h	60% ^e

^a With 5 mol % of Pd(TFA)₂, 20 mol % of pyridine, 2 equiv of Na₂CO₃, 500 mg/mmol of MS3Å, 1 atm O₂, toluene (0.1 M), 80 °C. ^b The starting material was used as a mixture of *E* and *Z* olefins. ^c Isolated with 7% of the aldehyde. ^d Isolated with 7% of an olefin isomer. ^e Isolated as a 5:2.3:1 mixture of **35**/olefin isomer/aldehyde.

excellent yields of spirocycles (**31**) and fused ring systems (**33**, **35**). The mode of oxidative reactivity—cyclization versus alcohol oxidation—appears dependent not only on the substrate (i.e., primary vs secondary alcohols) but also on the specific palladium source (cf. Uemura's work¹⁷).

Cyclizations of Carboxylic Acids and Derivatives. A range of carboxylic acid and carboxylic acid derivatives were synthesized and subjected to our optimized conditions. For some carboxylic acid derivatives (**36**, **38**, **40**, and **44**), the addition of an inorganic base was found to be unnecessary, and exposure simply to 5 mol % of Pd(TFA)₂, 20 mol % of pyridine, 500 mg of MS3Å/mmol substrate and 1 atm O₂ in toluene at 80 °C led to the oxidatively cyclized products

Table 5. Oxidative Cyclizations of Carboxylic Acids and Derivatives^a

entry	substrate	product	time	yield
1. ^b			8 h	90%
2. ^b			8 h	88%
3. ^b			4 h	82%
	38, 40	39, 41		
4. ^b			48 h	63% ^{c,d}
5.			48 h	62% ^e
6.			2 h	77% ^f
7.			12 h	86% ^{f,g}
	42	43		
	44	45		
	46	47		
	48	49		

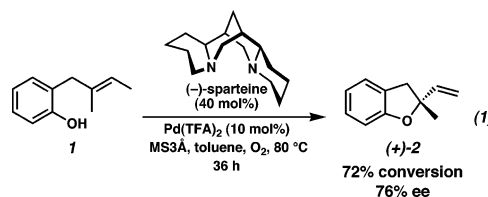
^a With 5 mol % of Pd(TFA)₂, 20 mol % of pyridine, MS3Å, 1 atm O₂, toluene, 80 °C. ^b The starting material was used as a mixture of *E*:*Z* olefins. ^c With 10 mol % of pyridine, 2 equiv of LiOAc. ^d *Z*:*E* 3:1. ^e With 10 mol % of Pd(TFA)₂, 40 mol % of pyridine. ^f With 10 mol % of Pd(TFA)₂, 40 mol % of pyridine, 2 equiv of Na₂CO₃. ^g Isolated with 6% of an olefin isomer.

(Table 5, entries 1–3, 5). Benzoic acids (**36**) and amides (**38**, **40**) are cyclized in good to excellent yields (entries 1–3). A β -keto ester (**42**) undergoes cyclization as a vinylogous acid to form a heterocycle (**43**) rather than a carbocycle (entry 4). Primary acid derivatives react to form spirocycles (**45**) or fused bicyclic systems (**47** and **49**), depending on the position of the olefin (entries 5–7). The cyclization of derivatives **46** and **48** is more facile with double the catalyst loading and is accelerated by the presence of 2 equiv of Na₂CO₃.

Reaction Scope and Limitations. The high yields, usually brief reaction times, and range of substrates that are characteristic of this aerobic Pd-catalyzed oxidative cyclization demonstrate the utility of the simple reaction conditions—Pd, ligand, base, O₂, and solvent. Nearly identical conditions are applicable to five different types of nucleophiles: phenols, primary alcohols, carboxylic acids, a vinylogous acid, and amides. Electron-rich phenols are excellent substrates, and multiple olefin substitution patterns are tolerated. Primary alcohols undergo oxidative cyclization without significant oxidation to the aldehyde, a fact that illustrates the range of reactivity available from various Pd(II) salts under differing conditions. In addition to the cyclization of a benzylic alcohol, nonbenzylic alcohols can form both fused and spirocyclic ring systems; the same is true of the acid derivatives. Undoubtedly the range of alcohol substrates could be increased.²² While phenol/olefin,

alcohol/olefin,^{5a,6a,23} and carboxylic acid/olefin^{5c,24,25} cyclizations have been achieved before under Pd(II)/oxidant catalysis, our conditions differentiate themselves by meeting the criteria for extension to an asymmetric version: simplicity, capacity to accommodate a chiral ligand, and active catalysis in a noncoordinating solvent. Without a system of this type, the development of a direct aerobic asymmetric cyclization would be less feasible.

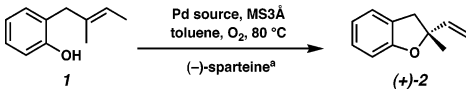
Elaboration of Racemic Conditions to an Asymmetric Version. A number of chiral ligands were screened with the conditions established for the racemic cyclizations, including typical chiral ligands, such as bisoxazolines, as well as less common frameworks for catalysis, such as brucine (not shown, see Supporting Information). As we observed during the development of our oxidative kinetic resolution chemistry,^{15a} the natural product (–)-sparteine was by far the most successful at inducing asymmetry in the cyclization. Treatment of **1** with Pd(TFA)₂ in the presence of (–)-sparteine, MS3Å, and O₂ in toluene led to a 72% conversion to 76% enantioenriched dihydrobenzofuran (+)-**2** after 36 h (eq 1).²⁶



Optimization of the Asymmetric Reaction: Screen of Palladium Sources. With (–)-sparteine as the most enantioselective candidate to emerge from the chiral ligand screen, we set out to increase selectivity through an optimization of Pd(II) source and basic additive. This seemed essential in light of what we had observed during the development of the racemic reaction (cf. Table 1), as well as with the kinetic resolution chemistry.^{15a,27} Pd(II) halide sources provided product in some cases, but with degradation of enantiomeric excess (Table 6, entries 1–2, 4–5). Pd(OAc)₂ is more effective at inducing asymmetry, but at the expense of conversion (entry 6). The extent of asymmetric induction varied surprisingly in the presence of different Pd(II) sources. For example, Pd(COD)Cl₂ results in only 10% ee (entry 2), whereas Pd(TFA)₂ provides 76% ee (entry 7). It is remarkable that this change has such a large effect on enantioselectivity.²⁸ Pd(TFA)₂ remained the optimal catalyst, and it was found that by using the complex of Pd(TFA)₂ and (–)-sparteine directly rather than generating it in situ led to slightly improved and more reliable results (entry 8). Generally, more electron-deficient Pd sources were more selective in the cyclization.

- (23) Semmelhack, M. F.; Kim, C. R.; Dobler, W.; Meier, M. *Tetrahedron Lett.* **1989**, 30, 4925–4928.
- (24) Åkermarck, B.; Larsson, E. M.; Oslob, J. D. *J. Org. Chem.* **1994**, 59, 5729–5733.
- (25) A similar cyclization of sulfonamides was recently reported: Fix, S. R.; Brice, J. L.; Stahl, S. S. *Angew. Chem., Int. Ed.* **2002**, 41, 164–166.
- (26) With 10 mol % of Pd(TFA)₂, 40 mol % of ligand, 500 mg/mmol MS3Å, 0.41 equiv of tridecane internal GC standard, 1 atm O₂, toluene (0.1 M), 80 °C. Conversion and enantiomeric excess were determined by GC.
- (27) Attempted cyclization under the chloroform-based rate-enhanced conditions for oxidative kinetic resolution ((sp)PdCl₂ (5 mol %), (–)-sparteine (7 mol %), Cs₂CO₃ (0.4 equiv), CHCl₃ (0.1 M), O₂ (1 atm, balloon), 25 °C) resulted in no reaction.
- (28) A similar, albeit less dramatic counterion effect was observed in the oxidative kinetic resolution of secondary alcohols; see ref 15a and Nielsen, R. J.; Keith, J. M.; Stoltz, B. M.; Goddard, W. A., III. *J. Am. Chem. Soc.* **2004**, 126, 7967–7974.

(22) For a recent review of Pd-catalyzed reactions of alcohols to form ether linkages, see: Muzart, J. *Tetrahedron* **2005**, 61, 5955–6008.

Table 6. Optimization of Pd Source and Additive for the Asymmetric Oxidative Cyclization


entry	Pd source	additive	time	conv ^b	ee ^b
1	Pd(nbd)Cl ₂	none	36 h	68%	12%
2	Pd(COD)Cl ₂	none	36 h	30%	10%
3	PdCl ₂	none	36 h	2%	12%
4	Pd(CH ₃ CN) ₂ Cl ₂	none	36 h	53%	12%
5	PdBr ₂	none	36 h	32%	8%
6	Pd(OAc) ₂	none	36 h	18%	51%
7	Pd(TFA) ₂	none	36 h	72%	76%
8 ^c	(sp)Pd(TFA) ₂ ^d	none	36 h	83%	77%
9	Pd(OTf) ₂ (CH ₃ CN) ₄	none	160 h	20%	16%
10 ^e	(sp)Pd(TFA) ₂	Na ₂ CO ₃	36 h	53%	76%
11 ^e	(sp)Pd(TFA) ₂	Ca(OH) ₂	36 h	87%	81%

^a With 10 mol % of Pd source, 40 mol % of (–)-sparteine, 500 mg/mmol of MS3A, 1 atm O₂, toluene (0.1 M), 80 °C. ^b Conversion measured by GC or by ¹H NMR. ^c With 30 mol % of (–)-sparteine. ^d sp = (–)-sparteine. ^e With 10 mol % of (sp)Pd(TFA)₂, 100 mol % of (–)-sparteine.

Table 7. Pd-Catalyzed Asymmetric Aerobic Oxidative Cyclization^a

entry	substrate	product	time	yield ^b	ee ^c
1.			36 h	87%	81%
2.			24 h 60 h @ 55 °C	64% 57%	88% 90%
3.			36 h	47%	83%
4.			36 h	47%	86%
5.			24 h	60%	20%

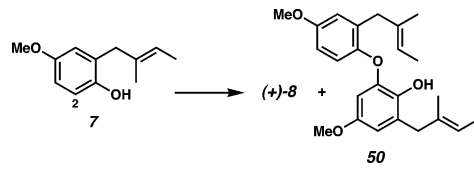
^a With 10 mol % of (sp)Pd(TFA)₂, 100 mol % of (–)-sparteine, 2 equiv of Ca(OH)₂, 500 mg/mmol of MS3A, 1 atm O₂, toluene (0.1 M), 80 °C. ^b Isolated yield. ^c Measured by GC.

However, switching the anion from trifluoroacetate to triflate resulted in degradation of the catalyst (formation of Pd black). In addition to palladium source, a basic additive affects reaction rate and selectivity. Best results occur with Ca(OH)₂, the presynthesized (sp)Pd(TFA)₂ complex, and 1 equiv of (–)-sparteine (entry 11).

Asymmetric Oxidative Cyclization of Phenol Substrates.

Under the optimized conditions, phenol **1** was cyclized to provide dihydrobenzofuran (+)-**2** in 81% ee and 87% isolated yield (Table 7, entry 1).²⁹ Application of these conditions to other substrates that reacted successfully under the racemic conditions proved difficult.³⁰ *p*-Methoxyphenol **7** is

(29) The absolute stereochemistry was determined by comparison of the optical rotation with that reported for (–)-**2**, by Uozumi and Hayashi, ref 12c (see Supporting Information for rotation data). We report the absolute stereochemistry of (+)-**8**, (–)-**6**, (+)-**4**, and (–)-**10** by analogy to this substrate.

Table 8. Aryl Ether Formation and Attempted Suppression


entry	additive	time	(+)- 8 : 50 ^b	ee of (+)- 8 ^c
1 ^a	amberlyst 15 resin (10 mol %)	24 h	5:3	87%
2	TsOH·H ₂ O (10 mol %)	24 h	3:1	81%
3	TsOH·H ₂ O (100 mol %)	80 h	6:1	50%
4	none	16 h	2:1	86%

^a With 10 mol % of (sp)Pd(TFA)₂, 100 mol % of (–)-sparteine, 500 mg/mmol of MS3A, 1 atm O₂, toluene (0.1 M), 80 °C. ^b Measured by ¹H NMR. ^c Measured by GC.

transformed with high selectivity to give (+)-**8** in 90% ee and 57% yield (entry 2). *tert*-Butylphenol **5** and *p*-methylphenol **3** do not react quickly but are obtained with good levels of enantiomeric excess (entries 3 and 4). *p*-Acylphenol **9** cyclizes, but with low percent enantiomeric excess, perhaps indicative of a change in mechanism for this less electron-rich substrate.

Unfortunately, the highly enantioenriched *p*-methoxy dihydrobenzofuran (+)-**8** is produced with a dimeric aryl ether byproduct (**50**, Table 8). This byproduct could form via palladation ortho to the phenol, followed by coupling to another molecule of substrate. The addition of various acids suppressed the formation of the byproduct, perhaps by decomposing the postulated palladium aryl species, but also depressed the enantioselectivity of the cyclization (entry 3).

While the asymmetric oxidative cyclization conditions that we have developed are not yet general, we have established that it is possible to adapt a direct dioxygen-coupled racemic reaction to aerobic asymmetric catalysis, which had not been achieved previously for this class of reaction. In conjunction with our work on the oxidative kinetic resolution of secondary alcohols, the asymmetric cyclization chemistry makes significant headway into the realm of enantioselective oxidase-type chemistry.

Mechanistic Investigations. Deuterium-Labeling Studies of Primary Alcohol Substrates in the Presence of a Monodentate Ligand. We hypothesized that the sluggishness of the asymmetric cyclization could be attributed to the difference in ligand denticity between pyridine and (–)-sparteine. In fact, in our screen of nitrogen-containing ligands (Table 2), the decreased reactivity of all the bidentate ligands confirmed that (–)-sparteine was not uniquely deactivating. We thus wished to gain an understanding of the mechanism of the cyclization, and whether a change from a mono- to a bidentate ligand would involve a change in mechanism. One commonly proposed mechanism for Pd(II)-catalyzed cyclizations involves activation of the olefin by Pd(II) followed by anti nucleophilic attack, that is, anti oxypalladation; indeed, this was our original hypothesis.^{18,31} The resulting Pd–alkyl intermediate would then undergo β-hydrogen elimination. Such a mechanism is remi-

(30) Tetrasubstituted olefin-containing phenol **21**, which is Hayashi's most selectively cyclized substrate,^{12a} fails to cyclize under our optimized conditions.

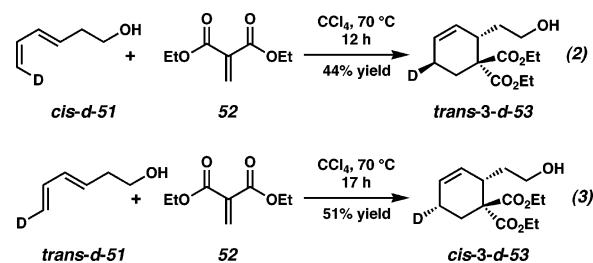
(31) (a) Ref 3g and references therein. (b) Coleman, J. P.; Hegedus, L. S. *Principles and Applications of Organometallic Chemistry*; University Science Books: Mill Valley, CA, 1980; pp 401–424.

niscent of that proposed by Bäckvall, Stille, and Kurosawa for the Wacker oxidation of ethylene to acetaldehyde.³² A second possible mechanism is supported by recent reports from Hayashi and Wolfe that find certain heteroatom/olefin cyclizations catalyzed by Pd(II) occur with syn oxypalladation.³³ Wolfe further suggests that this involves an olefin insertion into a Pd(II)–heteroatom bond.^{34–36} This pathway is analogous to the alternative mechanism for the Wacker oxidation proposed by Henry and others that involves syn oxypalladation.³⁷ A third mechanism entails allylic C–H activation by Pd(II) to form an intermediate π -allyl species that would then undergo reductive elimination with the heteroatom nucleophile.^{9,38,39}

We set out to differentiate these mechanisms by synthesizing stereospecifically deuterium-labeled substrates and observing the products of the cyclization in the presence of a mono- and bidentate ligand. Deuterium-labeled alcohols *trans*-**3-d-53** and *cis*-**3-d-53** were designed such that oxidative cyclization would result in retention or elimination of the label. Recently, during the course of this work, Hayashi and co-workers reported a similar study for the oxidative cyclization of olefin-appended phenols by Pd(II); in those examples, syn oxypalladation occurs exclusively except in the presence of additional chloride ion.^{33a} We chose to focus on primary alcohols because of the interesting dichotomy between oxidative cyclization and alcohol oxidation and because a study of this type had not yet been carried out for this substrate class.

Stereospecific deuterium incorporation into primary alcohol substrates could be effected with the Diels–Alder reaction shown in eqs 2 and 3 that completed the relatively straightforward synthesis of *trans*- and *cis*-**3-d-53**.⁴⁰ We assumed the constraints that a *cis* 6-5 fused ring system would form, and

that β -hydrogen elimination would only occur when the Pd atom and eliminated H or D atom are syn to each other.⁴¹ The presence or absence of a deuterium label in the product would then differentiate mechanistic pathways. Although a π -allyl mechanism could be difficult to unambiguously rule out with our test substrate, we hoped to at least distinguish syn from anti oxypalladation.⁴²

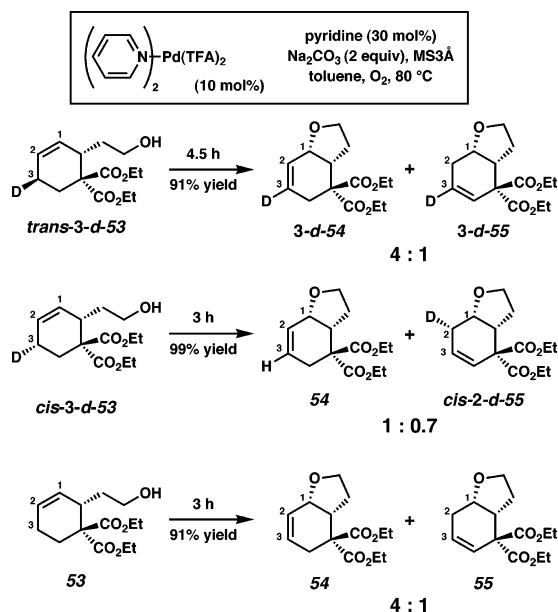


Treatment of *trans*-**3-d-53** with 10 mol % of (pyridine)₂Pd(TFA)₂, 20 mol % of pyridine, 2 equiv of Na₂CO₃, 1 atm O₂, and 500 mg/mmol of MS3Å in toluene at 80 °C for 3 h provided **3-d-54** along with olefin isomer **3-d-55** in a 4:1 ratio and 95% overall yield (Scheme 1). Likewise, reaction of the *cis* isomer (*cis*-**3-d-53**) under the same conditions led to the formation of a 1:0.7 mixture of undeuterated **54** and *cis*-**2-d-55** in nearly quantitative yield. Reaction of *cis*-**3-d-53** in the absence of Na₂CO₃ leads to a nearly identical result.⁴³ The stereochemistry of *cis*-**2-d-55** was confirmed by ¹H NMR homodecoupling and NOE experiments.⁴⁰ Reexposure of the product mixture (**54** and *cis*-**2-d-55**) to the same reaction conditions and an additional 10 mol % of starting material (*cis*-**3-d-53**) for 4 h resulted in an identical ratio of products in 90% isolated yield. Comparison of the ¹H and ²H NMR spectra of the products of the above reactions with those formed from undeuterated cyclohexene **53** confirmed the presence or absence of a deuterium label.⁴⁰

The mechanistic origin of the products, illustrated for the *cis* diastereomer (*cis*-**3-d-53**), is shown in Scheme 2. For comparison, all three possible pathways are shown. Path A involves anti nucleophilic attack of the Pd-coordinated olefin by the pendant alcohol or alkoxide. Subsequent β -hydrogen elimination would lead to the deuterium-labeled product **3-d-54**, but this product is not observed (vide supra, Scheme 1). In Path B, oxypalladation entails allylic C–H(D) activation and subsequent reductive elimination to Pd(0) upon formation of the C–O bond. The stereochemistry of the reductive elimination would likely be anti.⁴⁴ Further, unless selective C–D activation occurs, a mixture of labeled and unlabeled products would be expected; instead, a single product is observed. In Path C, a Pd–alkoxide undergoes syn oxypalladation followed by syn β -deuterium elimination to provide the observed major product, **54**. Reinsertion

- (32) (a) Bäckvall, J.-E.; Åkermark, B.; Ljunggren, S. O. *J. Chem. Soc., Chem. Commun.* **1977**, 264–265. (b) Bäckvall, J.-E.; Åkermark, B.; Ljunggren, S. O. *J. Am. Chem. Soc.* **1979**, *101*, 2411–2416. (c) James, D. E.; Hines, L. F.; Stille, J. K. *J. Am. Chem. Soc.* **1976**, *98*, 1806–1809. (d) Stille, J. K.; Divakaruni, R. *J. Am. Chem. Soc.* **1978**, *100*, 1303–1304. (e) Majima, T.; Kurosawa, H. *J. Chem. Soc., Chem. Commun.* **1977**, 610–611.
- (33) For systems containing oxygen nucleophiles, see: (a) Hayashi, T.; Yamasaki, K.; Mimura, M.; Uozumi, Y. *J. Am. Chem. Soc.* **2004**, *126*, 3036–3037. (b) Hay, M. B.; Hardin, A. R.; Wolfe, J. P. *J. Org. Chem.* **2005**, *70*, 3099–3107. For systems containing nitrogen nucleophiles, see: (c) Ney, J. E.; Wolfe, J. P. *Angew. Chem., Int. Ed.* **2004**, *43*, 3605–3608. (d) Ney, J. E.; Wolfe, J. P. *J. Am. Chem. Soc.* **2005**, *127*, 8644–8651.
- (34) Bäckvall and co-workers have characterized computationally the reactivity of nucleophiles toward *cis* migration, or olefin insertion, in (π -olefin)Pd(II) complexes and found that the HOMO–LUMO gap between the π^* olefin orbital and a Pd–OH nucleophile is too large for migration to be frontier controlled. Rather, the process is charge controlled, and likely does not occur through a concerted, four-center transition state as in true olefin insertion reactions of Pd–CH₃ or Pd–H nucleophiles. See: (a) Bäckvall, J.-E.; Björkman, E. E.; Pettersson, L.; Siegbahn, P. *J. Am. Chem. Soc.* **1984**, *106*, 4369–4373. (b) Bäckvall, J.-E.; Björkman, E. E.; Pettersson, L.; Siegbahn, P. *J. Am. Chem. Soc.* **1985**, *107*, 7265–7267.
- (35) For experimental evidence of insertion of tetrafluoroethylene into a Pt–O bond, see: Bryndza, H. E. *Organometallics* **1985**, *4*, 406–408.
- (36) For a recent example of olefin insertion into a rhodium amide, see: Zhao, P.; Krug, C.; Hartwig, J. F. *J. Am. Chem. Soc.* **2005**, *127*, 12066–12073.
- (37) (a) Francis, J. W.; Henry, P. M. *Organometallics* **1991**, *10*, 3498–3503. (b) Francis, J. W.; Henry, P. M. *Organometallics* **1992**, *11*, 2832–2836. (c) Zaw, K.; Henry, P. M. *Organometallics* **1992**, *11*, 2008–2015. (d) Hamed, O.; Henry, P. M. *Organometallics* **1997**, *16*, 4903–4909. (e) Hamed, O.; Thompson, C.; Henry, P. M. *J. Org. Chem.* **1997**, *62*, 7082–7083. (f) Hamed, O.; Henry, P. M.; Thompson, C. *J. Org. Chem.* **1999**, *64*, 7745–7750. (g) ten Brink, G.-J.; Arends, I. W. C. W.; Papadogiannakis, G.; Sheldon, R. A. *Appl. Catal., A* **2000**, *194–195*, 435–442. (h) Nelson, D. J.; Li, R.; Brammer, C. *J. Am. Chem. Soc.* **2001**, *123*, 1564–1568.
- (38) For evidence supporting this mechanism in intermolecular examples, see: (a) Grennberg, H.; Simon, V.; Bäckvall, J.-E. *J. Chem. Soc., Chem. Commun.* **1994**, 265–266. (b) Grennberg, H.; Bäckvall, J.-E. *Chem.–Eur. J.* **1998**, *4*, 1083–1089.
- (39) Trost has shown that Pd(TFA)₂ will form π -allyl complexes by C–H activation of olefins in acetone: Trost, B. M.; Metzner, P. J. *J. Am. Chem. Soc.* **1980**, *102*, 3572–3577.
- (40) See Supporting Information for details.

- (41) Examples of anti β -hydrogen elimination under various conditions have been reported, but involve aromatization or the formation of highly conjugated systems. See: (a) Toyota, M.; Ilangovan, A.; Okamoto, R.; Masaki, T.; Arakawa, M.; Ihara, M. *Org. Lett.* **2002**, *4*, 4293–4296. (b) Lautens, M.; Fang, Y.-Q. *Org. Lett.* **2003**, *5*, 3679–3682. (c) Hughes, C. C.; Trauner, D. *Angew. Chem., Int. Ed.* **2002**, *41*, 1569–1572. (d) Hennessy, E. J.; Buchwald, S. L. *J. Am. Chem. Soc.* **2003**, *125*, 12084–12085.
- (42) For a study that attempts to differentiate between oxypalladation and a π -allyl route for a Pd(II)-catalyzed cyclization, see: Zanon, G.; Porta, A.; Meriggi, A.; Franzini, M.; Vidari, G. *J. Org. Chem.* **2002**, *67*, 6064–6069.
- (43) After 5.5 h, 89% yield of a 1:0.75 mixture of **54** and *cis*-**2-d-55** was obtained.
- (44) For a Pd(II)– π -allyl electrophile, a primary alcohol (or alkoxide) would fall into the class of “soft” nucleophiles, the conjugate acids of which as defined by Trost have a $\text{p}K_{\text{a}} < 25$. See: Trost, B. M.; Van Vranken, D. L. *Chem. Rev.* **1996**, *96*, 395–422.

Scheme 1. Cyclization of Deuterium-Labeled Primary Alcohol Substrates with Pyridine as Ligand

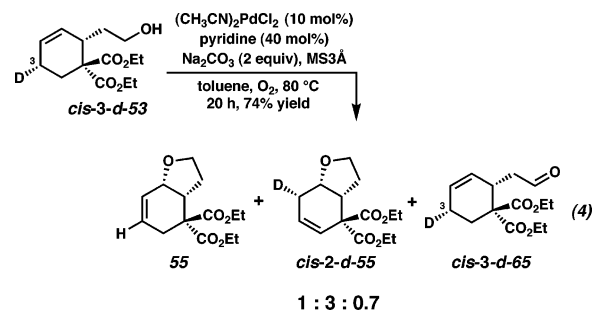
tion of the Pd–D intermediate to the product olefin accounts for formation of the observed minor product, *cis*-2-d-55.⁴⁵ The fact that an identical product mixture is obtained upon reexposure of the products to the reaction conditions supports the occurrence of reinsertion before dissociation of the product olefin from the Pd–D fragment.⁴⁶ Finally, because the Pd⁰ intermediate in Path B cannot easily account for the formation of olefin isomer *cis*-2-d-55, we favor syn oxypalladation Path C.

The remaining steps of the catalytic cycle involve reprocessing of Pd(II) by molecular oxygen. On the basis of earlier work by Murahashi and Takehira,^{3g,47} Uemura has proposed that this occurs by insertion of O₂ directly into the [Pd^{II}]-H to form a Pd–hydroperoxide intermediate.^{17a} Stahl has elaborated the details of a mechanism for aerobic oxidation catalysis by Pd(II) that entails reductive elimination of HX from [Pd^{II}]-H to form Pd(0), which is then oxidized by O₂ in the formation of a Pd–peroxo intermediate.^{48,49} In any case, these steps and those outlined in Scheme 2 proceed through Pd(II) and Pd(0) intermediates; however, we cannot definitively rule out a process involving Pd(IV).⁵⁰

We strove to further discount the π -allyl mechanism (Path B, Scheme 2) by comparing the reactivity of two phenol substrates. As shown in Scheme 3, a phenol with a disubstituted pendant olefin (**23**) cyclizes in good yield to provide dihydrobenzofuran **24**. The most likely π -allyl species to form from **23** would involve activation of a benzylic proton (**63**), but

intramolecular nucleophilic attack of this intermediate (**63**) would not lead to **24**. The π -allyl intermediate that could lead to cyclization (**64**) could be expected to arise from terminal olefin **62** as well (Scheme 3); both starting materials would lead to the same cyclized product (**24**). Instead, treatment of **62** with our standard racemic conditions leads to a complex mixture after several hours and less than 5% of **24**. It may be that allylic or benzylic C–H activation does occur, but is not productive for cyclization.

To explore whether the large effect of Pd(II) source on reactivity originated from a change in mechanism, the cyclization of *cis*-3-d-53 was carried out under the same reaction conditions but in the presence of (pyridine)₂PdCl₂ (eq 4).⁵¹ The initial cyclization product (**55**) was identical to that obtained under the conditions employing (pyridine)₂Pd(TFA)₂; however, the olefin isomer *cis*-2-d-55 was now the major product. The cyclic ethers were obtained along with aldehyde *cis*-3-d-65 in 74% overall yield after 20 h. The formation of ether **55** and ether *cis*-2-d-55 implies that syn oxypalladation still occurs, contrary to what Hayashi observed for a similar phenol substrate, for which anti oxypalladation dominates upon the addition of chloride ion.^{33a} The formation of aldehyde *cis*-3-d-65 likely occurs from a common alkoxide intermediate and highlights the effect that subtle changes in reaction conditions can have on the mode of oxidation.

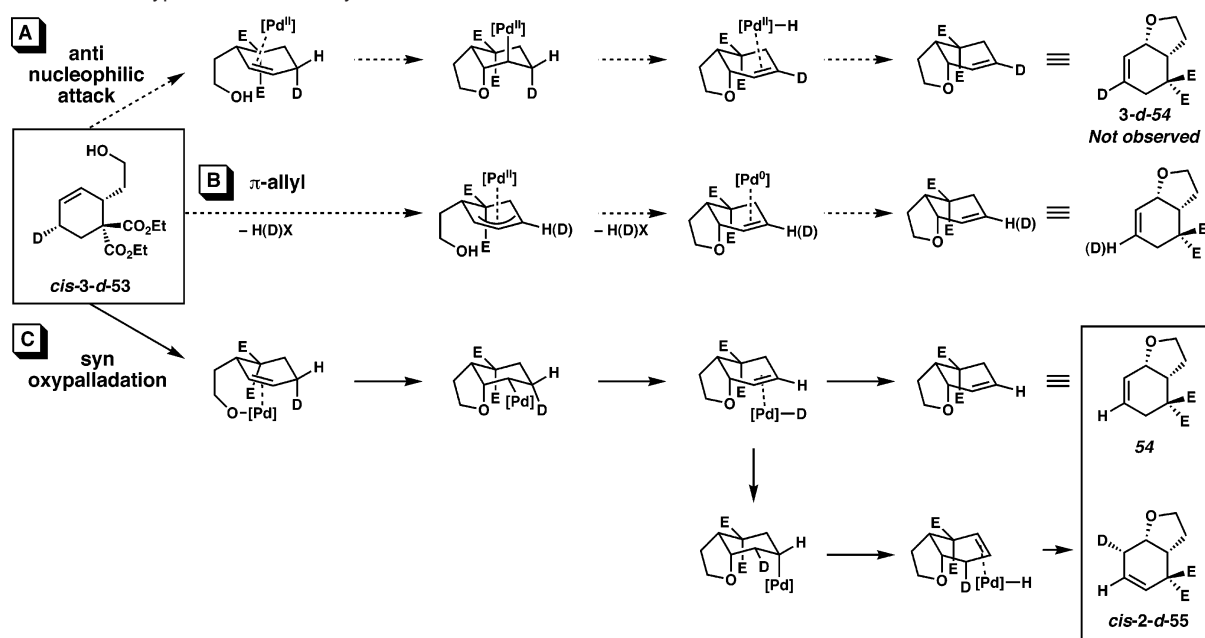


Deuterium-Labeling Studies of Primary Alcohol Substrates in the Presence of a Bidentate Ligand. The deuterium-labeled alcohol substrates *trans*-3-d-53 and *cis*-3-d-53 were then cyclized under conditions that employed dipyridyl as a bidentate ligand. Treatment of both *trans*-3-d-53 and *cis*-3-d-53 with 10 mol % of (dipyridyl)Pd(TFA)₂, Na₂CO₃ (2 equiv), 1 atm O₂, and 500 mg/mmol of MS3A in toluene at 80 °C led to the formation of the same major products (**3-d-54** and **54**, respectively) observed with the conditions that use pyridine (Scheme 4). None of the olefin isomer was observed; instead, a small amount of aldehyde (*trans*-3-d-65 and *cis*-3-d-65) was formed

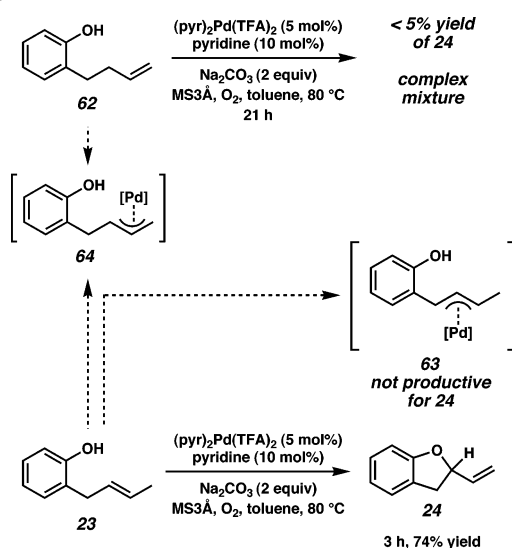
- (45) The difference in the product ratios formed from the *trans* and *cis* isomers could be accounted for by an isotope effect in the reinsertion step. Possibilities include slower reinsertion for the [Pd]–H bound olefin than for the [Pd]–D bound olefin, faster dissociation for the [Pd]–H bound olefin, faster reductive elimination of HX before dissociation or reinsertion, or faster coordination and insertion of O₂ into the [Pd]–H bond, if turnover of the catalyst occurs by that mechanism. At this time, we cannot rule out any of or distinguish among these possibilities.
- (46) Because both [Pd]–D and [Pd]–H are formed, reinsertion after [Pd] dissociates from product **54** would lead to scrambling of the isotopic label.
- (47) Takehira, K.; Hayakawa, T.; Orita, H. *Chem. Lett.* **1985**, 1835–1838 and references therein.
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- (50) In one scenario, oxidative activation of an allylic C–H bond could lead to a Pd(IV)–alkyl intermediate for the formation of a Pd(II) π -allyl species, or reductive elimination upon formation of the C–O bond could lead to a Pd(II) intermediate. In another scenario, Pd(II) could undergo oxidative activation of an O–H bond to form a Pd(IV)–alkoxide, although this is unlikely in the presence of base. Although such pathways cannot be ruled out, the recently reported Pd oxidation reactions proposed to involve Pd(IV) also involve strong oxidizing agents, such as PhI(OAc)₂ and I₂. For recent articles invoking Pd(IV) in Pd-catalyzed oxidative processes, see: (a) Dick, A. R.; Hull, K. L.; Sanford, M. S. *J. Am. Chem. Soc.* **2004**, *126*, 2300–2301. (b) Dick, A. R.; Kampf, J. W.; Sanford, M. S. *Organometallics* **2005**, *24*, 482–485. (c) Alexanian, E. J.; Lee, C.; Sorensen, E. J. *J. Am. Chem. Soc.* **2005**, *127*, 7690–7691. (d) Yoneyama, T.; Crabtree, R. H. *J. Mol. Catal. A* **1996**, *108*, 35–40. (e) For a related Pd-catalyzed iodination system, see: Giri, R.; Chen, X.; Yu, J.-Q. *Angew. Chem., Int. Ed.* **2005**, *44*, 2112–2115.
- (51) The catalyst was generated in situ by heating (CH₃CN)₂PdCl₂ and all other reagents except the substrate for 15 min before addition of the substrate.

Scheme 2. Three Oxypalladation Pathways



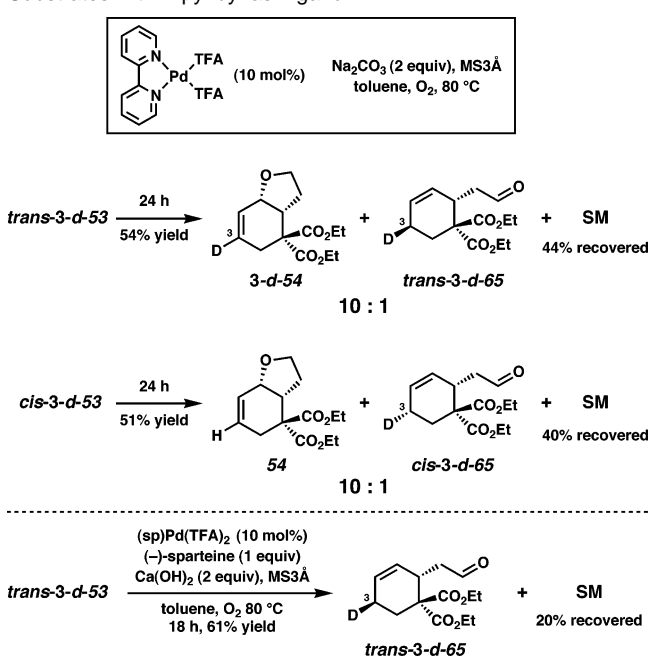
Scheme 3. Attempted Cyclization of a Terminal Olefin-Appended Phenol



for both starting material diastereomers.⁵² Both diastereomers were slower to cyclize in the presence of (dipyridyl)Pd(TFA)₂ than in the presence of (pyridine)₂Pd(TFA)₂. The reaction of *trans*-3-d-53 under the asymmetric conditions using (–)-sparteine led only to oxidation to the aldehyde. The major product distribution from cyclization in the presence of both a mono- and bidentate ligand is the same. Presumably, the aldehyde forms via β -hydrogen elimination from a common alkoxide intermediate before oxypalladation can occur.

On the basis of the results outlined above, we propose that, at least for the alcohol substrates, reaction under mono- or bidentate ligated conditions occurs by a similar pathway. The difference in reaction rate could arise from the degree to which each ligand can stabilize intermediates under otherwise identical reaction conditions. A neutral monodentate ligand such as

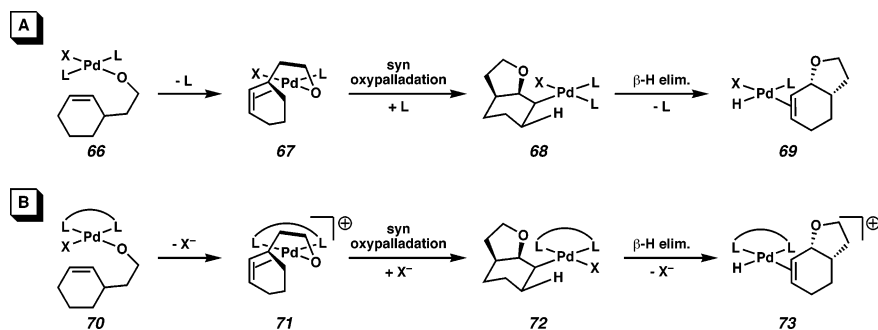
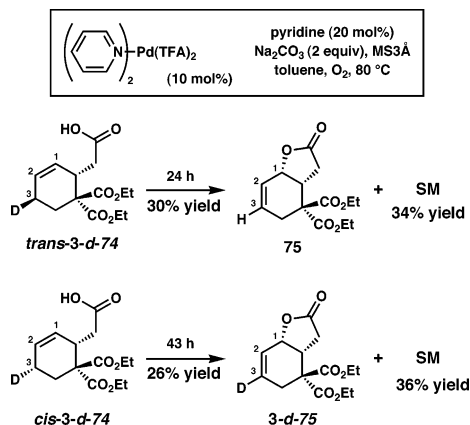
Scheme 4. Cyclization of Deuterium-Labeled Primary Alcohol Substrates with Dipyridyl as Ligand



pyridine can dissociate when necessary to free a coordination site for the substrate while maintaining charge-neutral intermediates (**67**–**69**, Scheme 5, A). When dipyridyl is used, neutral ligand dissociation may be more difficult due to chelation, and instead, an anionic ligand must dissociate, resulting in charged intermediates (**71** and **73**, Scheme 5, B). Such intermediates may be higher in energy under the reaction conditions, which in turn results in decreased reactivity.⁵³ (–)-Sparteine (cf. Scheme 4) not only leads to the same limitation but also prevents cyclization, perhaps due to steric congestion at Pd that prevents formation of a Pd–O–olefin chelate (**71**).

(53) We cannot rule out the possibility that the reaction with dipyridyl as ligand proceeds through similar intermediates with slow, partial dissociation of one of the dipyridyl nitrogen atoms. The reaction with (–)-sparteine as ligand is much less likely to undergo partial dissociation due to its structural rigidity.

(52) The identity of the aldehyde was confirmed by oxidation of the alcohol by another method; see Supporting Information.

Scheme 5. Comparison of Reaction Pathways with a Monodentate and Bidentate Ligand**Scheme 6.** Cyclization of Deuterium-Labeled Carboxylic Acid Substrates with Pyridine as Ligand

Although an extension of these mechanistic conclusions to the racemic and asymmetric phenol cyclizations may be tenuous, the same scenario would account for the slowness of the reactions catalyzed by ((-)-sparteine)Pd(TFA)₂ relative to (pyridine)₂Pd(TFA)₂. Hayashi and co-workers' results also support this conclusion. Their system undergoes syn oxypalladation in the presence of a bidentate ligand with a dicationic Pd(II) catalyst in methanol using benzoquinone as the terminal oxidant.^{33a} The higher dielectric solvent and dicationic catalyst may facilitate the formation of the charged intermediates shown in Scheme 5B, whereas our toluene-based conditions slow their formation. Because most of the chiral ligands that are effective in inducing asymmetry are bidentate, the mechanistic implication of the results described above is that selectivity *and* good reactivity will be difficult to attain in the Pd(II)/toluene/O₂ system.

Deuterium-Labeling Studies of Carboxylic Acid Substrates. Alcohols *trans*-3-*d*-53 and *cis*-3-*d*-53 were oxidized⁴⁰ to the corresponding carboxylic acid derivatives (*trans*-3-*d*-74 and *cis*-3-*d*-74) and subjected to the cyclization conditions that employ pyridine (Scheme 6). Both were slow to cyclize, and formation of Pd black was observed; nevertheless, some product was obtained and analyzed by ¹H and ²H NMR. In contrast to the alcohol substrates, the formation of unlabeled lactone **75** from *trans*-3-*d*-74 and labeled lactone **3-d-75** from *cis*-3-*d*-74 indicates that cyclization of the acid proceeds through anti oxypalladation. The olefin isomer arising from reinsertion was not observed. Cyclization of *trans*-3-*d*-74 in the absence of Na₂CO₃ resulted in greatly diminished yield (4% yield after 24 h) of the same product (**75**) but slower formation of Pd black, which indicates that the inorganic base does not affect the

stereochemistry of oxypalladation, but may play a role in catalyst decomposition.

Although the reasons for the change in product distribution at this point remain unclear, geometrical constraints, p*K*_a differences, or differences in nucleophilicity between an acid and an alcohol are all possibilities. The additional unsaturation in the forming lactone could make geometrically unfavorable the Pd–O–olefin chelate that would precede syn oxypalladation. Whatever the case, these results confirm the versatility and indiscriminating reactivity of Pd(II) catalysts in oxidation reactions. As demonstrated by these results and the accumulating evidence from several groups, generalizations about reactivity and mechanism that span different substrates and different reaction conditions are difficult to make for palladium-catalyzed oxidation reactions.

Summary and Conclusion

In summary, we have demonstrated that oxidase-type cyclizations of several different nucleophiles onto pendant olefins can be achieved in excellent yields under simple conditions. Reactivity is highly dependent on Pd(II) source, basic additive, and ligand. Several different types of cyclic systems are attainable, including aryl and alkyl bicycles, as well as fused and spirocyclic motifs. These cyclizations are part of an ongoing effort to develop enantioselective oxidase-type reactions using Pd(II) catalysis and molecular oxygen. To this end, the racemic conditions we developed were suitable for extension to an enantioselective cyclization that uses the chiral ligand (–)-sparteine. While the asymmetric oxidative cyclization conditions that we have developed are not yet general, we have established that it is possible to adapt a direct dioxygen-coupled racemic reaction to aerobic asymmetric catalysis, which had not before been achieved for this class of reaction. In conjunction with our work on the oxidative kinetic resolution of secondary alcohols, the asymmetric cyclization chemistry is a significant stride into the arena of enantioselective oxidase-type chemistry.

Stereospecifically deuterated primary alcohol substrates were used to probe the stereochemistry of oxypalladation and to gain insight into the mechanism of cyclization. Contrary to the common mechanistic proposal for reactions of this type (i.e., “Wacker cyclization” or anti oxypalladation), reaction of the primary alcohol substrates appears to occur via syn oxypalladation. This is in agreement with more recent reports from Hayashi and Wolfe on related systems. Neither the presence of chloride anion nor the use of a bidentate instead of a monodentate ligand changes the stereochemistry of oxypalladation. The implications for the asymmetric reaction are that bidentate ligands may destabilize the intermediates necessary for a syn

oxypalladation pathway; thus reactivity is decreased relative to the conditions employing a monodentate ligand. On the other hand, similarly deuterated carboxylic acid substrates react in the opposite sense, that is, via anti oxypalladation.

As contemporary work in this rapidly developing, recently reborn field continues to demonstrate, oxidase-type catalysis by Pd(II) is highly versatile and adaptable to a variety of applications.⁵⁴ Our studies reported herein emphasize the subtleties of reactivity and mechanism in this field. Further development of dehydrogenation reactions using Pd(II) and molecular oxygen, racemic and asymmetric, is ongoing.

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Supporting Information Available: Full experimental details and characterization data for all new compounds. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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