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Effective Molarity (EM): Quantitative measure of the ease of ring closure

For studies of reaction rates:  $EM = k_{intra}/k_{inter}$  (1)

For equilibrium studies:  $EM = K_{intra}/K_{inter}$  (2)

In case that's not enough math:  $EM = e \left[ -(\Delta H_{intra} - \Delta H_{inter})/RT \right] \cdot e \left[ (\Delta S_{intra} - \Delta S_{inter})/R \right]$ 

(3)

 $\Delta H$  and  $\Delta S$  should be read as  $\Delta H^{\circ}$  and  $\Delta S^{\circ}$  (equilibrium case) and  $\Delta H^{\ddagger}$  and  $\Delta S^{\ddagger}$  (rate case) (derived from applying either TS theory or thermodynamics to Eq. 1 and 2)

But we can make Eq. 3 easier to handle:  $EM = EM_H \times EM_S$  (4)

Bottom line: Enthalpy and Entropy independently affect the ease of ring closure

Galli, C.; Mandolini, L. Eur. J. Org. Chem. 2000, 3117-3125.

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The EM for a system represents an intramolecular reactivity that has been corrected for the inherent reactivity of the end groups

Just a few comments on Eq. 3:  $EM = e^{[-(\Delta H_{intra} - \Delta H_{inter})/RT]} \cdot e^{[(\Delta S_{intra} - \Delta S_{inter})/R]}$  (3)

 $(\Delta H^{\circ}_{intra} - \Delta H^{\circ}_{inter}) = strain energy of the ring$ 

 $(\Delta H^{\dagger}_{intra} - \Delta H^{\dagger}_{inter})$  = strain energy of the ring-shaped TS

 $(\Delta S_{intra} - \Delta S_{inter})$  depends solely on the number of skeletal bonds in the bifunctional precursor undergoing cyclization

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